

ATOMIC STRUCTURE QUIZ

Name: _____

Period: _____

Complete the following table

Element	Protons	Electrons	Neutrons	Mass #	Isotope Name/Symbol
Magnesium	_____	_____	_____	25	_____
_____	33	_____	42	_____	_____
_____	_____	10	_____	21	_____
_____	_____	_____	_____	_____	Potassium - 40
_____	_____	13	14	_____	_____

Complete the following table

Element	Mass #	Atomic #	Atomic Mass	Neutrons	Isotope Name/Symbol
Zinc	_____	_____	_____	35	_____
_____	52	23	_____	_____	_____
_____	_____	4	_____	5	_____
_____	_____	_____	_____	_____	Mercury - 200
Calcium	40	_____	_____	_____	_____

Answer the following questions:

1. What is the charge of a proton _____ neutron _____ electron _____
2. What particle(s) contribute the most mass to an atom _____
3. What particle identifies an element as that specific element _____
4. Two isotopes exist for Potassium. Potassium - 39 and Potassium - 40. Which isotope is the most abundant? _____
5. Calculate the average atomic mass of Magnesium. The three magnesium isotopes have an atomic mass and relative abundances of 23.985 amu (78.99%), 24.986 amu (10.00%), and 25.982 amu (11.01%). **SHOW ALL WORK AND UNITS BELOW!**